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What is a chemical Equation Class

10? Chemical Equation:

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Representation of chemical reaction using symbols and formulae of the substances is called Chemical Equation.

Example:  $A + B \rightarrow C + D$  In this equation, A and B are called reactants and C and D are called the products. The arrow shows the direction of the chemical reaction.

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The chapter of class 10 on chemical reactions and equations is therefore essential for helping students in forming a strong foundational base for higher classes. It explains in detail the elementary concepts of a chemical reaction and chemical

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Chemical Reactions and Equations Class 10 Important Questions Short Answer Type I It is a displacement reaction which is highly exothermic. The amount of heat evolved is so large that the metal is... The displacement reaction of iron (III) oxide with aluminium is known as thermite reaction.

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in short format

## Answers

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EXOTHERMIC REACTIONS □ An exothermic reaction is a chemical reaction that releases energy by light or heat. It is the opposite of an endothermic reaction. □

Expressed in a chemical equation: reactants → products + energy

For example 1)  $4\text{Fe(s)} + 3\text{O}_2\text{(g)} \rightarrow 2\text{Fe}_2\text{O}_3\text{(s)}$  2) Making ice cubes 3) Mixing sodium sulfite and  $\text{SO}_2$ .

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Chemical Equation: A mathematical statement which represents the chemical reaction.

Types of Chemical Reactions:

Combustion reaction – The reaction of a combustible

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material with an oxidiser to give an oxidised product.

Decomposition reaction – The reaction in which a single component breaks down into multiple products.

~~ICSE Selina Class 8 Chemistry Chapter 6 Chemical ...~~

Balanced chemical equation →  
The chemical equation that shows the chemical reaction needs to be balanced. A balanced chemical equation occurs when the number of the atoms involved in the reactants side is equal to the number of atoms in the products side.

$$\text{Zn} + \text{H}_2\text{SO}_4 \rightarrow \text{ZnSO}_4 + \text{H}_2$$
$$3\text{Fe (s)} + 4\text{H}_2\text{O (g)} \rightarrow \text{Fe}_3\text{O}_4 \text{ (s)} + 4\text{H}_2 \text{ (g)}$$

~~Notes of Ch 1 Chemical Reactions~~



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